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CHEMISTRY

( Major )

Paper : 6·2

( Physical Chemistry )

Full Marks : 60

Time : 3 hours

*The figures in the margin indicate full marks  
for the questions*

*(Symbols signify their usual meaning)*

1. Answer the following in brief : 1×7=7

- (a) Name the crystal system with characteristics  $a = b \neq c$ ;  $\alpha = \beta = 90^\circ$ ,  $\gamma = 120^\circ$ .
- (b) Write the Miller indices of the plane which intersects the  $x$ -axis at  $2a$  and which is parallel to  $y$ - and  $z$ -axes.
- (c) Write the definition of partition function.

- (d) Generally the mass average molar mass of a polymer is greater than its number average molar mass. State when they become equal.
- (e) Name the type of polymerization which results in Nylon-66.
- (f) State how doping enhances the electrical conductivity of a semiconductor.
- (g) State where the octahedral voids of the fcc unit cell are located.
2. (a) The edge length in NaCl crystal is  $5.63 \times 10^{-10}$  m. Find the distance between (111) planes. 2
- (b) Explain why alkali metals are soft. 2
- Or
- KCl acquires magenta colour when crystals of the compound is exposed to potassium vapours. Explain the observation. 2



- (c) Consider a system of 6 distinguishable particles. One of the macrostates of the system has the following distribution of particles :

Energy level	:	0	1	2	3	4
Number of particles	:	0	0	2	2	2

Find thermodynamic probability. 2

- (d) Explain how the formation of micelles affects the electrical conductivity of soap solution. 2

Or

Explain why protein-in-water sol undergoes coagulation on addition of alcohol. 2

3. Define systematic error and random error. State how these are related to precise and accurate measurements. An experiment was carried out to determine the amount of a metal in a sample and the result was found to be 35.68% while the true value is 35.98%. Find relative error. 2+1+2=5

Or

Define average deviation and standard deviation. Estimation of Fe present in a

sample showed the following results in a series of experiments :

<i>Experiment</i>	<i>Amount of Fe</i>
I	7.146%
II	7.098%
III	6.942%
IV	7.256%
V	6.593%

Find average deviation and standard deviation. 2+3=5

4. Answer either (a) and (b) or (c) and (d) : 5
- (a) Find the ratio between the populations of the two states indicated by I and II, such that the energy difference between state II and state I, ( $E_{II} - E_I$ ) is  $kT$ . The degeneracy in level I is 1 and that in level II is 3. 3
- (b) Calculate the internal energy of 1 mol He at 25 °C. 2
- (c) Calculate the characteristic vibrational temperature of  $O_2$  if its fundamental vibrational wave number is  $2337 \text{ cm}^{-1}$ . 3
- (d) Using Stirling approximation, find the value of  $\ln(100!)$ . 2



5. Answer either (a) or [(b) and (c)] : 5

(a) A mixture of two polymers contains  $w$  kg of each of the two. The molar mass of one polymer is  $10 \text{ kg mol}^{-1}$  and that of the other is  $20 \text{ kg mol}^{-1}$ . Calculate number average molar mass, mass average molar mass and polydispersity index.  $2+2+1=5$

(b) The osmotic pressure of  $1 \text{ m}^3$  of a solution containing  $2.5 \text{ kg}$  of a polymer is found to be  $250 \text{ Pa}$  at  $298 \text{ K}$ . Assuming that the solution does not deviate from ideal behaviour, calculate the molar mass of the polymer. 3

(c) Write the different steps through which addition polymerization occurs. 2

6. Answer either [(a), (b) and (c)] or [(d), (e) and (f)] : 10

(a) State Bragg's law and deduce the equation

$$2d \sin \theta = n\lambda \quad 4$$

(b) Superconductivity in metal is observed only by cooling it to near absolute zero. Explain this observation. 3

- (c) Show that the packing efficiency in ccp structure is nearly 74%. 3
- (d) In case of the ionic compounds of the type  $BA$ , explain how the radii of the cation and the anion influence packing of the smaller ion in different holes. 4
- (e) Write the difference between ferromagnetism and antiferromagnetism with respect to domain. 3
- (f) Using band theory, explain how electrical conductivity of conductor and semiconductor varies with temperature. 3

7. Answer [(a) and (b)] or [(c) and (d)] : 10

(a) Assuming a diatomic molecule to be rigid rotator, write the expression for rotational energy. Hence deduce an expression for the rotational partition function. What do you mean by characteristic rotational temperature?  
 $1+3+1=5$

(b) Using the concept of partition function, deduce an expression for the internal energy of monatomic ideal gas. Hence find an expression for the heat capacity at constant volume.  $3+2=5$



(c) A particle of mass  $m$  is moving inside a box of length  $a$ ,  $b$  and  $c$  along  $x$ -,  $y$ - and  $z$ -axes respectively. The potential inside the box is assumed to be zero. Find an expression for the translational partition function for a particle. 5

(d) Using partition function, deduce an expression for the entropy of monatomic gas. 5

8. Answer either [(a), (b) and (c)] or [(d) and (e)] : 10

(a) What do you mean by protection of colloid? Explain the mechanism of protection of colloid. 2+2=4

(b) Write about the processes responsible for the charge of colloidal particles. Discuss in brief how the co-ions and counter-ions are distributed around the charged colloidal particles. 2+2=4

(c)  $\text{Fe}(\text{OH})_3$  sol contains positively charged colloidal particles. In case of coagulation of this sol, the flocculation value of  $\text{K}_3[\text{Fe}(\text{CN})_6]$  is  $0.096 \text{ millimol L}^{-1}$  while that of  $\text{K}_2\text{SO}_4$  is  $0.210 \text{ millimol L}^{-1}$ . Explain the reason behind this difference in flocculation value. 2

(d) Discuss how the molecular mass of polymer can be determined by measuring osmotic pressure of its solution.

5

(e) What do you mean by condensation polymerization? Discuss about the kinetics of this type of polymerization.

1+4=5

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